6-78 A hot water stream is mixed with a cold water stream. For a specified mixture temperature, the mass flow rate of cold water is to be determined.

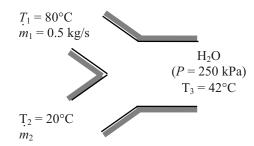
Assumptions 1 Steady operating conditions exist. 2 The mixing chamber is well-insulated so that heat loss to the surroundings is negligible. 3 Changes in the kinetic and potential energies of fluid streams are negligible. 4 Fluid properties are constant. 5 There are no work interactions.

Properties Noting that $T < T_{sat @ 250 \text{ kPa}} = 127.41^{\circ}\text{C}$, the water in all three streams exists as a compressed liquid, which can be approximated as a saturated liquid at the given temperature. Thus,

$$h_1 \cong h_{f@.80^{\circ}C} = 335.02 \text{ kJ/kg}$$

 $h_2 \cong h_{f@.20^{\circ}C} = 83.915 \text{ kJ/kg}$
 $h_3 \cong h_{f@.42^{\circ}C} = 175.90 \text{ kJ/kg}$

Analysis We take the mixing chamber as the system, which is a control volume. The mass and energy balances for this steady-flow system can be expressed in the rate form as



Mass balance:

$$\dot{m}_{\rm in} - \dot{m}_{\rm out} = \Delta \dot{m}_{\rm system}^{70 \text{ (steady)}} = 0 \longrightarrow \dot{m}_1 + \dot{m}_2 = \dot{m}_3$$

Energy balance:

$$\begin{split} \underline{\dot{E}_{\rm in} - \dot{E}_{\rm out}}_{\rm Rate\ of\ net\ energy\ transfer} &= \underbrace{\Delta \dot{E}_{\rm system}}_{\rm Rate\ of\ change\ in\ internal,\ kinetic,} = 0 \\ \dot{E}_{\rm in} &= \dot{E}_{\rm out} \\ \dot{m}_1 h_1 + \dot{m}_2 h_2 &= \dot{m}_3 h_3 \quad (\text{since}\ \dot{Q} = \dot{W} = \Delta \text{ke} \cong \Delta \text{pe} \cong 0) \end{split}$$

Combining the two relations and solving for \dot{m}_2 gives

$$\dot{m}_1 h_1 + \dot{m}_2 h_2 = (\dot{m}_1 + \dot{m}_2) h_3$$
$$\dot{m}_2 = \frac{h_1 - h_3}{h_3 - h_2} \dot{m}_1$$

Substituting, the mass flow rate of cold water stream is determined to be

$$\dot{m}_2 = \frac{(335.02 - 175.90) \text{ kJ/kg}}{(175.90 - 83.915) \text{ kJ/kg}} (0.5 \text{ kg/s}) = \mathbf{0.865} \text{ kg/s}$$